

Q1

Name the following binary ionic compounds:

a) $\text{Al}_2\text{O}_3(\text{s})$ b) $\text{HgI}_2(\text{s})$ c) $\text{Na}_3\text{P}(\text{s})$ d) $\text{CaBr}_2(\text{s})$

Q4

Write the chemical formula for the following binary ionic compounds:

a) zinc oxide c) magnesium iodide

b) iron(II) sulfide d) cobalt(III) chloride

Q5

Name the following binary covalently bonded compounds:

a) $\text{SF}_6(\text{g})$ b) $\text{N}_2\text{O}_5(\text{s})$ c) $\text{PCl}_5(\text{s})$ d) $\text{CF}_4(\text{g})$

Q6

Write the chemical formula for the following covalently bonded compounds:

a) dihydrogen monoxide c) dinitrogen tetroxide

b) sulfur trioxide d) dinitrogen monoxide

Q7

Name the following polyatomic ionic compounds:

a) $\text{K}_3\text{PO}_4(\text{s})$ b) $\text{NH}_4\text{Cl}(\text{s})$ c) $\text{LiClO}_4(\text{s})$ d) $\text{NaHCO}_3(\text{s})$

Q8

Write the chemical formula for the following polyatomic ionic compounds:

a) potassium hypochlorite c) sodium cyanide

b) magnesium oxalate d) ammonium sulfate

Supplemental:

Additional Practice, Frequently Asked in Exams

Name the following compounds:

HCl

Fe_2O_3

$\text{Au}(\text{ClO}_4)_3$

NH_4Cl

N_2O_3

NaOH

B_2Br_4

$(\text{NH}_4)_2\text{S}$

$\text{Sr}(\text{H}_2\text{PO}_4)_2$

BrF_5

Cl_2

H_2SO_4

PbO_2

Na_2SO_4

KCN

$\text{Fe}(\text{NO}_3)_3$

H_3PO_4

$\text{Mn}(\text{ClO}_2)_2$

$(\text{NH}_4)_2\text{CrO}_4$

Na_2S

H_2S

V_3N_5

$\text{Mg}(\text{OH})_2$

HCO_3

HOCl

Al_2O_3

HgI_2